



TEST

"Practice Test (1B-Periods)



10 Questions



10 min

Topics

Atomic Size, Ionization Energy and Electron Affinity,
Melting and Boiling Point, Electrical Conductivity
+ Metallic and non-metallic character, Oxidation
States + Hydration Energy, Halides & Oxides

Start Test

09 : 57



1/10



10 min



Hint

Q : For which molecule the bonding pair of electrons is equally shared between the atoms?



HF



HCl



H₂



BF₃

1

2

3

4

5

6

7

09 : 55



2/10



10 min



Hint

Q : Which group of periodic table shown generally abnormal trends of ionization energy.



IIA and VA



IIIA and IVA



IIIA and VIA



VIIIA

1

2

3

4

5

6

7

09 : 53



3/10



10 min



Hint

Q : Which of the followings are isoelectronic ions?



O^{2-}, F^{-}



Na^{+}, F^{-}



Na^{+}, Al^{+3}



All of these

1

2

3

4

5

6

7

09 : 51



4/10



10 min



Hint

Q : Which of the following elements forms acidic oxide only



Na



Al



Mg



Br

1

2

3

4

5

6

7

09 : 42



5/10



10 min



Hint

Q : Choose the amphoteric oxide



Na_2O



BaO



Al_2O_3



SO_3

1

2

3

4

5

6

7

09 : 39



6/10



10 min



Hint

Q : The oxidation number which is not shown by halogens



-1



+1



+2



+7

1

2

3

4

5

6

7

09 : 37



7/10



10 min



Hint

Q : Which of the following shows minimum electrical conductance



Na



Al



Mg



Cu

1

2

3

4

5

6

7

09 : 34



8/10



10 min



Hint

Q : Lanthanide contraction is related to decrease in



Atomic number



Atomic mass



Atomic energy



Atomic size

4

5

6

7

8

9

10

09 : 33



9/10



10 min



Hint

Q : The electronic configurations of some elements are given below. The element with highest electron affinity is



A $1s^2, 2s^2, 2p^6, 3s^2, 3p^5$



B $1s^2, 2s^2, 2p^6, 3s^2, 3p^1$



C $1s^2, 2s^2, 2p^5$



D $1s^2, 2s^2, 2p^2$

09 : 30



10/10



10 min



Hint

Q : The correct atomic size order is



Li > Mg



Be > Al



B > Si



Ba > Na

4

5

6

7

8

9

← "Practice Test (



Correct



Unattempted



Incorrect



1/10

Q : For which molecule the bonding pair of electrons is equally shared between the atoms?



HF



HCl



H₂



BF₃

← "Practice Test (



Correct



Unattempted



Incorrect



2/10

Q : Which group of periodic table shown generally abnormal trends of ionization energy.



IIA and VA



IIIA and IVA



IIIA and VIA



VIIIA

← "Practice Test (



Correct



Unattempted



Incorrect



3/10

Q : Which of the followings are isoelectronic ions?



$\text{O}^{2-}, \text{F}^{-}$



$\text{Na}^{+}, \text{F}^{-}$



$\text{Na}^{+}, \text{Al}^{+3}$



All of these



"Practice Test (



Correct



Unattempted



Incorrect



4/10

Q : Which of the following elements forms acidic oxide only



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Al



Mg



Br



"Practice Test (



Correct



Unattempted



Incorrect



5/10

Q : Choose the amphoteric oxide



Na_2O



BaO



Al_2O_3



SO_3



"Practice Test (



Correct



Unattempted



Incorrect



6/10

Q : The oxidation number which is not shown by halogens



-1



+1



+2



+7

1

2

3

4

5

6

7



"Practice Test (



Correct



Unattempted



Incorrect



7/10

Q : Which of the following shows minimum electrical conductance



Na



Al



Mg



Cu

← "Practice Test (



Correct



Unattempted



Incorrect



8/10

Q : Lanthanide contraction is related to decrease in



Atomic number



Atomic mass



Atomic energy



Atomic size



Correct



Unattempted



Incorrect



9/10

Q : The electronic configurations of some elements are given below. The element with highest electron affinity is



$1s^2, 2s^2, 2p^6, 3s^2, 3p^5$



$1s^2, 2s^2, 2p^6, 3s^2, 3p^1$



$1s^2, 2s^2, 2p^5$



$1s^2, 2s^2, 2p^2$



"Practice Test (



Correct



Unattempted



Incorrect



10/10

Q : The correct atomic size order is



Li > Mg



Be > Al



B > Si



Ba > Na



TEST

Test Level-1 (1B-Periods)



20 Questions



20 min

Topics

Explain the basic functions of the digestive system and the role of the digestive system in the body. Identify the organs of the digestive system and the role of the digestive system in the body. Identify the organs of the digestive system and the role of the digestive system in the body.

Start Test

19 : 56



1/20

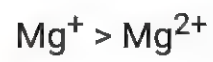


20 min



Hint

Q : Select the correct option according to size



All



2

3

4

5

6

7

19 : 54



2/20



20 min



Hint

Q : Which factor does not effect ionization energy across the period



Number of shell



Nuclear Charge



Nature of orbital



Both A & C

1



3

4

5

6

7

19 : 52



3/20



20 min



Hint

Q:

Mark the correct statement?



Metallic character increases down the group



Metallic character increases along a period



Metallic character decreases down the group



Metallic character remains the same down the group

1

2



4

5

6

7

19 : 49



4/20



20 min



Hint

Q :

Which is the pair of metalloids



Na and K



F and Cl



As and Sb



Cu and Au

1

2

3



5

6

7

19 : 47



5/20



20 min



Hint

Q : Not a polymeric halide



BeCl_2



CCl_4



AlCl_3



GaCl_3

1

2

3

4



6

7

19 : 45



6/20



20 min



Hint

Q : Which of the following is correct order with respect to covalent nature



$\text{PCl}_5 > \text{PCl}_3$



$\text{MgCl}_2 > \text{AlCl}_3$



$\text{PCl}_3 > \text{PCl}_5$



All are correct orders

1

2

3

4

5



7

19 : 44



7/20



20 min



Hint

Q : For elements of group VA, the oxidation state is equal to



The number of electrons present in the valance shell



The number of vacancies available in the valence shell



The number of all electrons present in all shells



Both A and B

1

2

3

4

5

6

19 : 40



8/20



20 min



Hint

Q : Which of following is a correct order of degree of hydration in alkali metal ions



6

7



9

10

11

12

19 : 38



9/20



20 min



Hint

Q : Keeping the size of atom in view, which order is correct?



$Mg > Sr$



$Ba > Mg$



$Lu > Ce$



$Cl > I$

6

7

8



10

11

12

19 : 35



10/20



20 min



Hint

Q : which of the following groups show abnormal trends in electron affinity values



IIA



VA



VIIIA



All of these

6

7

8

9

10

11

12

19 : 34



11/20



20 min



Hint

Q : Which of the following will have zero oxidation state?



IA



VIA



VIIA



VIIIA

6

7

8

9

10

11

12

19 : 31



12/20



20 min



Hint

Q : Electrical conductance will be high for



Metals



Semi-metals



None-metals



Coinage metals

6

7

8

9

10

11



19 : 27



13/20



20 min



Hint

Q : Elements form halides with halogens. Which of these do not form polymeric halide



Cs



Be



Al



Ga

10

11

12



14

15

16

17

19 : 25



14/20



20 min



Hint

Q : Which of the following shows minimum electrical conductance?



Na



Al



Mg



Cu

10

11

12

13



15

16

17

19 : 23



15/20



20 min



Hint

Q : Which ion will have the maximum value of hydration energy



Na^+



Cs^+



Ba^{2+}



Ca^{2+}

19 : 21



16/20



20 min



Hint

Q : The elements for which the value of electron affinity is high



Gain electrons readily



Gain electrons with difficulty



Lose electrons less readily



Lose electron readily



19 : 19



17/20



20 min



Hint

Q : The correct sequence of ionization energy of the elements is



S < P < Si < Al



Al < Si < P < S



Al < Si < S < P



P < S < Al < Si

11

12

13

14

15

16



19 : 16



18/20



20 min



Hint

Q : Which of the following series of elements is listed in order of decreasing atomic radius



N, S, F



Li, Na, K



Na, Mg, Al



F, Br, Cl

14

15

16

17



19

20

19 : 14



19/20



20 min



Hint

Q :

Melting point of I-A and II-A group elements decreases down the group due to _____.



Strong electronegativity



Strong attractive forces



Increment in size



High ionization energy

14

15

16

17

18



20

19 : 12



20/20

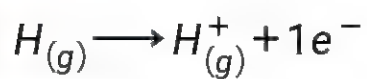
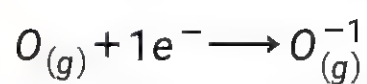
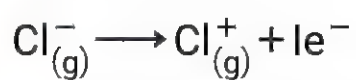
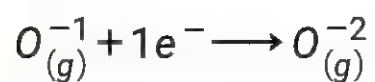


20 min



Hint

Q : Which of the following process is exothermic





TEST RESULT

Test Level-1 (1B-Periods)



20



20 min



08-Aug-2020



0 sec

Result Detail



	Correct	0
	Incorrect	0
	Unattempted	20

← Test Level-1 (1B-



Correct



Unattempted



Incorrect



1/20

Q : Select the correct option according to size



$\text{Mg}^+ > \text{Mg}^{2+}$



$\text{Ca}^+ > \text{Ca}^{+2}$



$\text{Be}^{+2} < \text{Be}^{+1}$



All



Correct



Unattempted



Incorrect



2/20

Q : Which factor does not effect ionization energy across the period



Number of shell



Nuclear Charge



Nature of orbital



Both A & C



Test Level-1 (1B-



Correct



Unattempted



Incorrect



3/20

Q:

Mark the correct statement?



Metallic character increases down the group



Metallic character increases along a period



Metallic character decreases down the group



Metallic character remains the same down the group



Test Level-1 (1B-



Correct



Unattempted



Incorrect



4/20

Q:

Which is the pair of metalloids



Na and K



F and Cl



As and Sb



Cu and Au



Test Level-1 (1B-



Correct



Unattempted



Incorrect



5/20

Q : Not a polymeric halide



BeCl_2



CCl_4



AlCl_3



GaCl_3



Correct



Unattempted



Incorrect



6/20

Q : Which of the following is correct order with respect to covalent nature



$\text{PCl}_5 > \text{PCl}_3$



$\text{MgCl}_2 > \text{AlCl}_3$



$\text{PCl}_3 > \text{PCl}_5$



All are correct orders



Test Level-1 (1B-



Correct



Unattempted



Incorrect



7/20

Q : For elements of group VA, the oxidation state is equal to



The number of electrons present in the valance shell



The number of vacancies available in the valence shell



The number of all electrons present in all shells



Both A and B



Correct



Unattempted



Incorrect



8/20

Q : Which of following is a correct order of degree of hydration in alkali metal ions





Test Level-1 (1B-



Correct



Unattempted



Incorrect



9/20

Q : Keeping the size of atom in view, which order is correct?



$Mg > Sr$



$Ba > Mg$



$Lu > Ce$



$Cl > I$



Test Level-1 (1B-



Correct



Unattempted



Incorrect



10/20

Q : which of the following groups show abnormal trends in electron affinity values



IIA



VA



VIIIA



All of these



Correct



Unattempted



Incorrect



11/20

Q : Which of the following will have zero oxidation state?



IA



VIA



VIIA



VIIIA

← Test Level-1 (1B-



Correct



Unattempted



Incorrect



12/20

Q : Electrical conductance will be high for



Metals



Semi-metals



None-metals



Coinage metals

← Test Level-1 (1B-



Correct



Unattempted



Incorrect



13/20

Q : Elements form halides with halogens. Which of these do not form polymeric halide



Cs



Be



Al



Ga

← Test Level-1 (1B-



Correct



Unattempted



Incorrect



14/20

Q : Which of the following shows minimum electrical conductance?



Na



Al



Mg



Cu



Correct



Unattempted



Incorrect



15/20

Q : Which ion will have the maximum value of hydration energy



Na⁺



Cs⁺



Ba²⁺



Ca²⁺



Test Level-1 (1B-



Correct



Unattempted



Incorrect



16/20

Q : The elements for which the value of electron affinity is high



A Gain electrons readily



B Gain electrons with difficulty



C Lose electrons less readily



D Lose electron readily



Correct



Unattempted



Incorrect



17/20

Q : The correct sequence of ionization energy of the elements is



$S < P < Si < Al$



$Al < Si < P < S$



$Al < Si < S < P$



$P < S < Al < Si$



Correct



Unattempted



Incorrect



18/20

Q : Which of the following series of elements is listed in order of decreasing atomic radius



N, S, F



Li, Na, K



Na, Mg, Al



F, Br, Cl



Correct



Unattempted



Incorrect



19/20

Q :

Melting point of I-A and II-A group elements decreases down the group due to _____.



Strong electronegativity



Strong attractive forces



Increment in size



High ionization energy



Correct



Unattempted

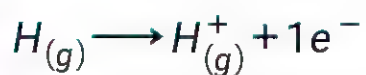
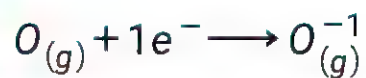
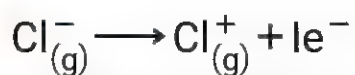
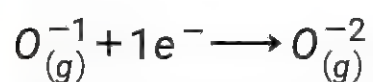


Incorrect



20/20

Q : Which of the following process is exothermic





TEST

Test Level-2 (Topic 1B)



30 Questions



25 min

Topics

Explain the basic functions of the digestive system and the role of the digestive system in the body. Identify the organs of the digestive system and the role of each organ. Explain the process of digestion and the role of the digestive system in the body. Identify the organs of the digestive system and the role of each organ. Explain the process of digestion and the role of the digestive system in the body.

Start Test

24 : 57



1/30



25 min



Hint

Q : Which of the following has minimum hydration energy



Li^+



Ca^{++}



Na^+



K^+



2

3

4

5

6

7

24 : 55



2/30



25 min



Hint

Q : Which of the following oxide is classified incorrectly



Na_2O — Neutral



CO_2 — Acidic



NO_2 — Acidic



BeO — Amphoteric

1



3

4

5

6

7

24 : 53



3/30



25 min



Hint

Q : Which of the following electronic configuration represents atoms of element having the highest 2nd ionization energy



$1s^2 2s^2 2p^4$



$1s^2 2s^2 2p^5$



$1s^2 2s^2 2p^6 3s^1$



$1s^2 2s^2 2p^6 3s^2$

1

2



4

5

6

7

24 : 51



4/30



25 min



Hint

Q : Which of the following is most acidic



Mn_2O_7



Mn_2O_3



MnO_2



MnO

1

2

3



5

6

7

24 : 49



5/30



25 min



Hint

Q : Which element has highest melting point



Be



C



B



N

1

2

3

4



6

7

24 : 48



6/30



25 min



Hint

Q : The p-block element among the following which is smallest in size is



Li



P



Na



Cl

1

2

3

4

5



7

24 : 46



7/30



25 min



Hint

Q : Which of the following does not form polymeric halide



Cs



Be



Al



Ga

1

2

3

4

5

6



24 : 43



8/30



25 min



Hint

Q : The increasing order of the first ionization enthalpies of the elements P, S and Cl is



Cl < S < P



P < S < Cl



P < Cl < S



S < P < Cl

5

6

7



9

10

11

24 : 41



9/30



25 min



Hint

Q : The most ionic chloride is



NaCl



CsCl



BaCl₂



MgCl₂



5

6

7

8



10

11

24 : 39



10/30



25 min



Hint

Q : The element whose normal oxide is amphoteric in nature



B



Al



Mg



C

5

6

7

8

9



11

24 : 37



11/30



25 min



Hint

Q : The melting and boiling point along 2nd and 3rd period



Increases from left to right



Decrease from left to right



1st increases upto group IVA and then decreases



1st decreases upto group IV A and then increases

5

6

7

8

9

10



24 : 34



12/30



25 min



Hint

Q : The ionization energy of boron is less than that of beryllium because



Atomic size of B > Be



Shielding effect of B < Be



Atomic size of B < Be



Removal of e^- from p-subshell

24 : 32



13/30

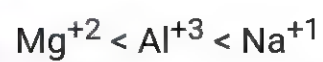
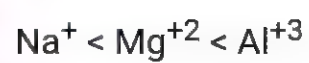
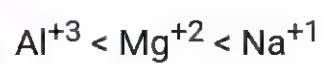
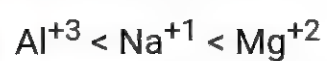


25 min



Hint

Q : Which among the following is the correct order of increasing ionic radius



24 : 30



14/30



25 min



Hint

Q : Which of the following shows minimum electrical conductance



Na



Al



Mg



Cu

24 : 28



15/30



25 min



Hint

Q : The ionization energy of nitrogen is more than that of oxygen because of



More attraction of electrons by the nucleus



Higher nuclear charge



The extra stability of half-filled p – orbitals



The size of nitrogen atom is smaller

24 : 25



16/30



25 min



Hint

Q :

An atom has electronic configuration: $1s^2, 2s^2, 2p^6, 3s^2, 3p^4$ You will place it in



Group II A , Period 2



Group VI A , Period 2



Group II A , Period 3



Group VI A , Period 3

24 : 24



17/30



25 min



Hint

Q : One of the following factors that has no effect on the 1st ionization energy along the 2nd and 3rd periods



Shielding effect



Nuclear charge



Half filled orbitals



Completely filled orbitals

1

12

13

14

15

16



18

24 : 21



18/30



25 min



Hint

Q : Which of the following is more covalent in nature



3

14

15

16

17

18

19

20

24 : 19



19/30



25 min



Hint

Q :

Nitrogen oxygen and fluorine have low melting points in 2nd period because



A They are small size elements



B They are non-metals



C They exist as individual molecules



D They are non-conductors

24 : 16



20/30



25 min



Hint

Q : Generally electronegativity of element in periods _____ and in groups _____



Increases, Increases



Increases, Decreases



Decreases, Decreases



Decreases, Increases

24 : 14



21/30



25 min



Hint

Q : Mg forms Mg^{+2} but does not form Mg^{+3} because of



Very high 2nd I.E



Very high 3rd I.E



Very low 2nd I.E



Very low 3rd I.E



24 : 13



22/30

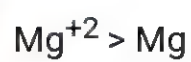
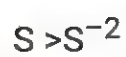


25 min



Hint

Q : The correct order of size is



24 : 11



23/30



25 min



Hint

Q : The ionic nature of 2nd period chlorides



Decreases



Increases



Remains constant



First increases and then decreases



24 : 09



24/30



25 min



Hint

Q : Among the following outermost electronic configuration of the least electronegative element in a given period is



$ns^2 np^5$



$ns^2 np^2$



$ns^2 np^4$



$ns^2 np^3$



24 : 07



25/30



25 min



Hint

Q : In which of the following pairs, the ionization energy of first specie is less than that of second



B, Be



N, O



N, P



P, S



24 : 05



26/30



25 min



Hint

Q : The ionic radii (\AA) of N^{3-} , O^{2-} and F^- are respectively



1.36, 1.40 and 1.71



1.36, 1.71 and 1.40



1.71, 1.40 and 1.36



1.71, 1.36 and 1.40



24 : 03



27/30



25 min



Hint

Q : Which of the following oxides is unlikely to dissolve in aqueous sodium hydroxide



SiO_2



MgO



Al_2O_3



P_4O_{10}



24 : 01



28/30



25 min



Hint

Q : Heat of hydration of gaseous hydrogen ion when it dissolved in water to make infinitely dilute solution is



-1075kJmole⁻¹



-1280kJmole⁻¹



-1175kJmole⁻¹



-965kJmole⁻¹

23 : 58



29/30



25 min



Hint

Q :

Melting and boiling points for first three elements of 3rd period are in order of $\text{Na} < \text{Mg} < \text{Al}$ It is due to



Increase of charge on metal ions



Increase in numbers of delocalized electrons



Increase in the strength of the metallic bonding



All of these

24

25

26

27

28



30

23 : 56



30/30



25 min



Hint

Q : Sum of number of all elements belonging to 2nd and 3rd periods collectively are



10



16



26



36

24

25

26

27

28

29





TEST RESULT

Test Level-2 (Topic 1B)



30



25 min



08-Aug-2020



0 sec

Result Detail



	Correct	0
	Incorrect	0
	Unattempted	30



Test Level-2 (To



Correct



Unattempted



Incorrect



1/30

Q : Which of the following has minimum hydration energy



Li⁺



Ca⁺⁺



Na⁺



K⁺



Correct



Unattempted



Incorrect



2/30

Q : Which of the following oxide is classified incorrectly



Na_2O — Neutral



CO_2 — Acidic



NO_2 — Acidic



BeO — Amphoteric

← Test Level-2 (To



Correct



Unattempted



Incorrect



3/30

Q : Which of the following electronic configuration represents atoms of element having the highest 2nd ionization energy



$1s^2 2s^2 2p^4$



$1s^2 2s^2 2p^5$



$1s^2 2s^2 2p^6 3s^1$



$1s^2 2s^2 2p^6 3s^2$

← Test Level-2 (To



Correct



Unattempted



Incorrect



4/30

Q : Which of the following is most acidic



Mn_2O_7



Mn_2O_3



MnO_2



MnO

← Test Level-2 (To



Correct



Unattempted



Incorrect



5/30

Q : Which element has highest melting point



Be



C



B



N



Correct



Unattempted



Incorrect



6/30

Q : The p-block element among the following which is smallest in size is



Li



P



Na



Cl

← Test Level-2 (To



Correct



Unattempted



Incorrect



7/30

Q : Which of the following does not form polymeric halide



Cs



Be



Al



Ga



Correct



Unattempted



Incorrect



8/30

Q : The increasing order of the first ionization enthalpies of the elements P, S and Cl is



Cl < S < P



P < S < Cl



P < Cl < S



S < P < Cl



Test Level-2 (To



Correct



Unattempted



Incorrect



9/30

Q : The most ionic chloride is



NaCl



CsCl



BaCl₂



MgCl₂



Test Level-2 (To



Correct



Unattempted



Incorrect



10/30

Q : The element whose normal oxide is amphoteric in nature



B



A



C



D



Correct



Unattempted



Incorrect



11/30

Q : The melting and boiling point along 2nd and 3rd period



Increases from left to right



Decrease from left to right



1st increases upto group IVA and then decreases



1st decreases upto group IV A and then increases



Correct



Unattempted



Incorrect



12/30

Q : The ionization energy of boron is less than that of beryllium because



Atomic size of B > Be



Shielding effect of B < Be



Atomic size of B < Be



Removal of e^- from p-subshell



Correct



Unattempted

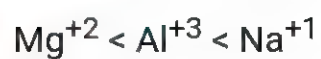
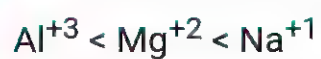
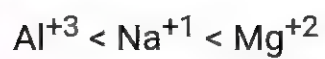


Incorrect



13/30

Q : Which among the following is the correct order of increasing ionic radius



← Test Level-2 (To



Correct



Unattempted



Incorrect



14/30

Q : Which of the following shows minimum electrical conductance



Na



Al



Mg



Cu



Test Level-2 (To



Correct



Unattempted



Incorrect



15/30

Q : The ionization energy of nitrogen is more than that of oxygen because of



More attraction of electrons by the nucleus



Higher nuclear charge



The extra stability of half-filled p – orbitals



The size of nitrogen atom is smaller



Correct



Unattempted



Incorrect



16/30

Q :

An atom has electronic configuration: $1s^2, 2s^2, 2p^6, 3s^2, 3p^4$ You will place it in



Group II A , Period 2



Group VI A , Period 2



Group II A , Period 3



Group VI A , Period 3



Correct



Unattempted



Incorrect



17/30

Q : One of the following factors that has no effect on the 1st ionization energy along the 2nd and 3rd periods



Shielding effect



Nuclear charge



Half filled orbitals



Completely filled orbitals



Test Level-2 (To



Correct



Unattempted



Incorrect



18/30

Q : Which of the following is more covalent in nature





Test Level-2 (To



Correct



Unattempted



Incorrect



19/30

Q :

Nitrogen oxygen and fluorine have low melting points in 2nd period because



A They are small size elements



B They are non-metals



C They exist as individual molecules



D They are non-conductors



Correct



Unattempted



Incorrect



20/30

Q : Generally electronegativity of element in periods _____ and in groups _____



Increases, Increases



Increases, Decreases



Decreases, Decreases



Decreases, Increases



Correct



Unattempted



Incorrect



21/30

Q : Mg forms Mg^{+2} but does not form Mg^{+3} because of



Very high 2nd I.E



Very high 3rd I.E



Very low 2nd I.E



Very low 3rd I.E



Correct



Unattempted

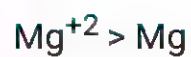
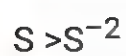


Incorrect



22/30

Q : The correct order of size is





Correct



Unattempted



Incorrect



23/30

Q : The ionic nature of 2nd period chlorides



Decreases



Increases



Remains constant



First increases and then decreases



Correct



Unattempted



Incorrect



24/30

Q : Among the following outermost electronic configuration of the least electronegative element in a given period is



$ns^2 np^5$



$ns^2 np^2$



$ns^2 np^4$



$ns^2 np^3$



Correct



Unattempted



Incorrect



25/30

Q : In which of the following pairs, the ionization energy of first specie is less than that of second



B, Be



N, O



N, P



P, S



Test Level-2 (To



Correct



Unattempted



Incorrect



26/30

Q : The ionic radii (\AA) of N^{3-} , O^{2-} and F^- are respectively



1.36, 1.40 and 1.71



1.36, 1.71 and 1.40



1.71, 1.40 and 1.36



1.71, 1.36 and 1.40



Test Level-2 (To



Correct



Unattempted



Incorrect



27/30

Q : Which of the following oxides is unlikely to dissolve in aqueous sodium hydroxide



SiO_2



MgO



Al_2O_3



P_4O_{10}

24

25

26

27

28

29

30



Correct



Unattempted



Incorrect



28/30

Q : Heat of hydration of gaseous hydrogen ion when it dissolved in water to make infinitely dilute solution is



-1075kJmole⁻¹



-1280kJmole⁻¹



-1175kJmole⁻¹



-965kJmole⁻¹



Correct



Unattempted



Incorrect



29/30

Q :

Melting and boiling points for first three elements of 3rd period are in order of $\text{Na} < \text{Mg} < \text{Al}$ It is due to



Increase of charge on metal ions



Increase in numbers of delocalized electrons



Increase in the strength of the metallic bonding



All of these



Test Level-2 (To



Correct



Unattempted



Incorrect



30/30

Q : Sum of number of all elements belonging to 2nd and 3rd periods collectively are



10



16



26



36